

(radiji atomov in ionov, ionizacijske energije, elektronska afiniteta).

Kemijska vez: ionska vez; kovalentna vez (nepolarna, polarna vez, dipolni moment, teorija valenčne vezi: principi teorije, resonanca, hibridizacija, geometrija molekul; teorija molekularskih orbital: principi teorije, delokalizirane MO); elektronegativnost; strukture anorganskih molekul (strukturne formule in nomenklatura anorganskih spojin)

Agregatna stanja snovi in medmolekulske vezi: plini, tekočine, trdne snovi; interakcije med molekulami (Van der Waalsove in vodikove vezi, vpliv teh vezi na lastnosti snovi). Struktura amorfnih in kristaliničnih trdnih snovi: ionski, kovalentni, kovinski in molekularski kristali, polprevodniki, tekoči kristali.

Disperzni sistemi: prave in koloidne raztopine ter njihove lastnosti.

Kemijske reakcije: kemijske reakcije in kemijske enačbe; energijske spremembe pri kemijskih reakcijah (standardne tvorbenne in standardne reakcijske entalpije, Hessov zakon); ravnotežje kemijskih reakcij, Le Chatelierovo načelo; vplivi na hitrost kemijske reakcije, kataliza; ionske reakcije (ionska ravnotežja, topnost, topnostni produkt); protolitske reakcije (Brønstedova teorija kislina in baz, pH, indikatorji, titracija, vpliv ionov na protolitska ravnotežja); redoks reakcije (oksidacijsko število in urejanje redoks reakcij, galvanski členi, elektroliza).

Koordinacijske spojine: stereokemijske značilnosti koordinacijskih spojin (izomerija); kemijska vez v koordinacijskih spojinah; vpliv elektronske konfiguracije na magnetne in optične lastnosti koordinacijskih ionov (kvalitativno).

Chemical bonding: ionic bond, covalent bond (non-polar, polar, dipolar momentum, valence bond theory: principles, resonance, hybridization, molecular geometry; molecular orbital theory: principles, delocalized MO); electronegativity, structures of inorganic compounds (structural formulas and nomenclature of inorganic compounds).

States of matter and intermolecular bonds: gases, liquids, solids, intermolecular interactions (Van der Waals and hydrogen bonds, influence of these bonds on properties of matter). Structure of amorphous and crystalline compounds: ionic, covalent and molecular crystals, semiconductors, liquid crystals.

Disperse systems: true and colloidal solutions and their properties.

Chemical reactions: chemical reactions and chemical equations: energy changes at chemical reactions (standard enthalpies of formation and standard reaction enthalpies, Hess law); chemical equilibrium, Le Chatelier's principle; the influences on the rate of the chemical reactions, catalysis; ionic reactions (ionic equilibria, solubility, solubility product); protolytic reactions (Brønsted theory of acids and bases, pH, indicators, titration. Influence of ions on protolytic equilibria). Redox reactions (oxidation number and balancing of redox reactions, galvanic cells, electrolysis).

Basics of coordination chemistry: stereochemical properties of coordination compounds, chemical bond in coordination compounds, the influence of the electronic structure on the magnetic and optical properties of coordination ions (qualitative basis).

Temeljna literatura in viri / Readings:

Osnovni učbenik:

- Boris Čeh, Splošna kemija, Založba UL FKKT, Ljubljana 2018, 389 str.

Dodatna literatura:

- R.H. Petrucci, W.S. Harwood, F.G. Herring, General Chemistry, Principles and modern applications, osma izdaja, Prentice Hall New Jersey, 2002, 1150 str. (40%) glede na interes študenta

- Erwin Riedel, Allgemeine und Anorganische Chemie, osma izdaja, Walter de Gruyter, Berlin, 2004, 380 str. (60%) glede na interes študenta

Cilji in kompetence:

Cilji: Poglobiti in nadgraditi znanje splošne in anorganske kemije, pridobljeno na srednji šoli, ki je potrebno za nadaljnji študij. Poudarek je na poznavanju in pravilnim razumevanjem osnovnih kemijskih zakonitosti ter poznavanju zgradbe snovi in njenega vpliva na kemijske lastnosti snovi.

Kompetence: Študent pozna in razume osnovne kemijske zakonitosti, principe in teorije ter jih zna uporabiti pri reševanju preprostih problemov (kvalitativno ali kvantitativno). Je sposoben poiskati in ovrednotiti določene podatke o snoveh in jih zna povezati z njihovimi lastnostmi.

Objectives and Competences:

Objectives: Deepening and upgrading the knowledge of general and inorganic chemistry, acquired in the secondary school, which is necessary for further study. Emphasis is given to knowledge and correct understanding basic chemical principles and knowledge on the constitution of matter and its influence on chemical properties of matter.

Competences: student knows and understands basic chemical principles and theories and knows how to use them for solving simple problems (qualitative or quantitative). He is able to find and evaluate given data about substances and is able to relate them to the properties of the substances.

Predvideni študijski rezultati:Znanje in razumevanje

Študent pozna in razume osnovne kemijske zakonitosti ter jih zna povezati z zgradbo in lastnostmi snovi in kemijskimi reakcijami.

Uporaba

Znanje in razumevanje osnovnih kemijskih zakonitosti so temelji predmetom pri nadaljnjem študiju.

Refleksija

Študent je sposoben oceniti pomen osnovnih kemijskih zakonitosti in teoretskega znanja za razlago eksperimentalnih dejstev in lastnosti snovi.

Prenosljive spretnosti

Študent zna poiskati podatke iz strokovne literature, podatke iz virov medmrežja pa zna kritično oceniti; zna uporabljati strokovni jezik (pisno in ustno).

Intended Learning Outcomes:Knowledge and Comprehension

Student knows and understands basic chemical principles and is able to relate them to the structure and properties of matter and chemical reactions.

Application

Knowledge and understanding of basic chemical principles are the basis of subjects for further study.

Analysis

Student is able to assess the meaning of basic chemical principles and theoretical knowledge for an explanation of experimental facts and properties of compounds.

Skill-transference Ability

Student is able to find data from professional literature and is able to critically evaluate the data from the internet; he is able to use the professional language (written and spoken).

Metode poučevanja in učenja:

Eksperimentalna predavanja z uporabo IKT; seminarji: sodelovalno učenje/ poučevanje ter problemsko delo; sprotno preverjanje znanja s testi.

Learning and Teaching Methods:

Experimental lectures using the ICT; seminars: cooperative learning/teaching and problem work; regular knowledge assessment using tests.

Načini ocenjevanja:

2 testa za sprotno preverjanje znanja in pisni izpit. Če študent na vsakem od obeh testov zbere najmanj 51 % točk je lahko oproščen opravljanja izpita. Ocenjevalna lestvica v skladu z enotno lestvico na Univerzi v Ljubljani: 6 – 10 opravi izpit, 1 – 5 ni opravi izpita.

Delež (v %) /

Weight (in %) /

Assessment:

2 test for during the semester and written exam. If the student collects at least 51 % of points at each of the tests, he can be excused from the exam. Grades according to the standard levels of the University of Ljubljana: 6-10 passed, 1-5 insufficient.

Reference nosilca / Lecturer's references:

- ZUPANIČ, Franc, MARKOLI, Boštjan, NAGLIČ, Iztok, WEINGÄRTNER, Tobias, **MEDEN, Anton**, BONČINA, Tonica. Phases in the Al-corner of the Al-Mn-Be system. *Microscopy and microanalysis*, ISSN 1431-9276. [Print ed.], FirstView Article, online: 18 June 2013, doi: [10.1017/S1431927613001852](https://doi.org/10.1017/S1431927613001852). [COBISS.SI-ID [16956694](https://www.cobiss.si/id/16956694)]

- IPAVEC, Andrej, GABROVŠEK, Roman, VUK, Tomaž, KAUČIČ, Venčeslav, MAČEK, Jadran, **MEDEN, Anton**. Carboaluminate phases formation during the hydration of calcite-containing Portland cement. *Journal of the American Ceramic Society*, ISSN 0002-7820, 2011, vol. 94, no. 3, str. 1238-1242, doi: [10.1111/j.1551-2916.2010.04201.x](https://doi.org/10.1111/j.1551-2916.2010.04201.x). [COBISS.SI-ID [34764037](https://www.cobiss.si/id/34764037)]

- MALI, Gregor, **MEDEN, Anton**, DOMINKO, Robert. [sup] 6 Li MAS NMR spectroscopy and first-principles calculations as a combined tool for the investigation of Li [sub] 2 MnSiO [sub] 4 polymorphs. *Chemical communications*, ISSN 1359-7345, 2010, issue 19, str.3306-8, doi: [10.1039/c003065a](https://doi.org/10.1039/c003065a). [COBISS.SI-ID [4386074](https://www.cobiss.si/id/4386074)]

1. P. Živec, F. Perdih, **I. Turel**, G. Giester, G. Psomas, Different types of copper complexes with the quinolone antimicrobial drugs ofloxacin and norfloxacin: Structure, DNA- and albumin-binding, *J. Inorg. Biochem.*, 117, 35–47 (2012).

2. D. Čurman, P. Živec, I. Leban, **I. Turel**, A. Polishchuk, K. D. Klika, E. Karaseva, V. Karasev, Spectral properties of Eu(III) compounds with antibacterial agent ciprofloxacin (cfqH). Crystal structure of [Eu(cfqH)(cfq)(H₂O)₄]Cl₂·4.55 H₂O, *Polyhedron*, 27, 1489-1496 (2008).

3. P. Drevenšek, J. Košmrlj, G. Giester, T. Skauge, E. Sletten, K. Sepčič, **I. Turel**, X-Ray Crystallographic, NMR and Antimicrobial Activity Studies of Magnesium Complexes of Fluoroquinolones - Racemic Ofloxacin and Its S-form, Levofloxacin, *J. Inorg. Biochem.*, 100, 1755-1763 (2006).